

**16****SOLUTIONS****SECTION 16.1 PROPERTIES OF SOLUTIONS (pages 471–477)**

This section describes the factors that affect the solubility of a substance and determines the rate at which it occurs dissolves.

**► Solution Formation (pages 471–472)**

Look at Figure 16.1 on page 471 to help you answer Questions 1 and 2.

- Underline the conditions that cause sugar to dissolve faster in water.
  - as a whole cube                      or                      as granulated sugar
  - when allowed to melt                      or                      when not melted
  - at a higher temperature                      or                      at lower temperature
- Name three factors that influence the rate at which a solute dissolves in a solvent.
  - \_\_\_\_\_
  - \_\_\_\_\_
  - \_\_\_\_\_
- In the following sentence, place an **all** or **few** particles dissolve more rapidly than larger particles because their particles expose a greater surface area to the colliding solvent molecules. \_\_\_\_\_

**► Solubility (pages 472–473)**

- Complete the following table, showing the steps in a procedure to determine the total amount of sodium chloride that will dissolve in 100 g of water at 25°C.

| Procedure                                   | Amount Dissolved | Amount Not Dissolved |
|---------------------------------------------|------------------|----------------------|
| Add 35.0 g of sodium chloride to the water. | 35.0 g           | 0.0 g                |
| Add an additional 1.0 g of sodium chloride. |                  |                      |
| Determine the total amount that dissolved.  |                  |                      |

- The amount of a substance that dissolves in a given quantity of solvent at a constant temperature is called the substance's \_\_\_\_\_ at that temperature.
- If a solution contains the maximum amount of solute for a given quantity of solvent at a constant temperature, it is called a(n) \_\_\_\_\_ solution.

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