

Unit 1 Progress Check: MCQ

1. Which of the following numerical expressions gives the number of moles in **5.0 g** of **CuO**?

A $5.0 \text{ g} \times 56 \text{ g/mol}$

B $\frac{5.0 \text{ g}}{56 \text{ g/mol}}$ ✓

C $56 \text{ g/mol} \times 5.0 \text{ g}$

D $\frac{5.0}{56} \text{ g} \times 56 \text{ g/mol}$

2. In a lab a student is given a **2.0 g** sample of pure **Cu** metal. Which of the following pieces of information is most useful for determining the number of **Cu** atoms in the sample? Assume that the pressure and temperature in the lab are **1.0 atm** and **25°C**.

A The molar mass of Cu ✓

B The density of Cu at 25°C

C The volume of the Cu sample

D The ratio of the two main isotopes found in pure Cu

3. A **1.0 mol** sample of which of the following compounds has the greatest mass?



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J Ma

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