



Name: _____

Date: _____

Student Exploration: Phase Changes (4b)

Directions: Follow the instructions to go through the simulation. Respond to the questions and prompts in the orange boxes.

Vocabulary: altitude, boil, boiling point, freeze, freezing point, gas, liquid, melt, melting point, phase, solid

Activity B: Temperature and molecular motion	Get the Gizmo ready: <ul style="list-style-type: none">Click Reset, and select the Micro view.Set Ice volume to 0 cm³.Set Add/remove heat energy to 0 J/s.	
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Question: Why do phase changes occur?

- Compare:** Set the **Water temperature** to **0 °C** and click **Play**. Observe the water molecules. Click **Reset**, set the **Water temperature** to **100 °C**, and click **Play** again.

What do you notice?

There is no difference.

- Observe:** Click **Reset**. The **mean molecular speed** of the water molecules is displayed below the container. Set the **Water temperature** to **0 °C** and **Add/remove heat energy** to **400 J/s**. Click **Play**.

- How does the mean speed of the water molecules change as they are heated?

The water gets hotter (has more heat energy).

- Does the mean molecular speed change as much as the temperature as the water heats up? Explain.

Yes, the more the temperature increases the more the molecules of the water speed up (until they hit the boiling point, where they cannot get hotter any longer.)

- Explain:** How is temperature related to the motions of molecules?

The higher the temperature, the faster the molecules move.

- Observe:** Click **Reset**. Set the **Water temperature** to **20 °C** and the **Ice volume** to **50 cm³**. Set **Add/remove heat energy** to **0 J/s**. Click **Play**. How do the molecules in the liquid interact with the molecules in the solid?

Phase Changes Gizmo Answer Sheet

Jin-Ying Zhang

Phase Changes Gizmo Answer Sheet:

New Scientist ,2006 *Phase Changes at Extremes* Jonathan Curley,2010

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