

## the Aufbau principle



**Atomic orbitals**  
When an electron occupies an orbital it has a definite quantum state.

**Pauli Exclusion Principle**  
Max 2 electrons in 1 orbital

**Quantum numbers**  
Quantum numbers define the properties of an electron, including its location.

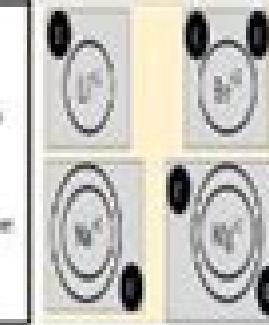
**Electron shells**  
Max 2 electrons in 1 shell

Atomic number	Element	Electron configuration
1	Hydrogen	1s <sup>1</sup>
2	Helium	1s <sup>2</sup>
3	Lithium	1s <sup>2</sup> 2s <sup>1</sup>
4	Boron	1s <sup>2</sup> 2s <sup>2</sup> 2p <sup>1</sup>
5	Carbon	1s <sup>2</sup> 2s <sup>2</sup> 2p <sup>2</sup>
6	Nitrogen	1s <sup>2</sup> 2s <sup>2</sup> 2p <sup>3</sup>
7	Oxygen	1s <sup>2</sup> 2s <sup>2</sup> 2p <sup>4</sup>
8	Fluorine	1s <sup>2</sup> 2s <sup>2</sup> 2p <sup>5</sup>
9	Neon	1s <sup>2</sup> 2s <sup>2</sup> 2p <sup>6</sup>
10	Chlorine	1s <sup>2</sup> 2s <sup>2</sup> 2p <sup>5</sup> 3s <sup>1</sup>
11	Sodium	1s <sup>2</sup> 2s <sup>2</sup> 2p <sup>6</sup> 3s <sup>1</sup>
12	Magnesium	1s <sup>2</sup> 2s <sup>2</sup> 2p <sup>6</sup> 3s <sup>2</sup>
13	Aluminum	1s <sup>2</sup> 2s <sup>2</sup> 2p <sup>6</sup> 3s <sup>2</sup> 3p <sup>1</sup>
14	Silicon	1s <sup>2</sup> 2s <sup>2</sup> 2p <sup>6</sup> 3s <sup>2</sup> 3p <sup>2</sup>
15	Phosphorus	1s <sup>2</sup> 2s <sup>2</sup> 2p <sup>6</sup> 3s <sup>2</sup> 3p <sup>3</sup>
16	Sulfur	1s <sup>2</sup> 2s <sup>2</sup> 2p <sup>6</sup> 3s <sup>2</sup> 3p <sup>4</sup>
17	Chlorine	1s <sup>2</sup> 2s <sup>2</sup> 2p <sup>6</sup> 3s <sup>2</sup> 3p <sup>5</sup>
18	Argon	1s <sup>2</sup> 2s <sup>2</sup> 2p <sup>6</sup> 3s <sup>2</sup> 3p <sup>6</sup>

**Electron arrangement**  
Electrons between Nucleus and Nucleus

**Periods**

1	Hydrogen
2	Helium
3	Lithium
4	Boron
5	Carbon
6	Nitrogen
7	Oxygen
8	Fluorine
9	Neon
10	Chlorine
11	Sodium
12	Magnesium
13	Aluminum
14	Silicon
15	Phosphorus
16	Sulfur
17	Chlorine
18	Argon

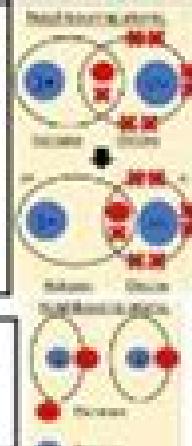


**Electronegativity**  
It is a measure of how strongly atoms attract the electrons in the bond it forms.

Chlorine has the shared electrons more because it contains more protons.

Chlorine has a slight negative charge due to having more electrons closer to it.

Hydrogen has a slight positive charge due to having fewer electrons and further away.



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CALCULATE REACTANT CONCENTRATIONS AND PRODUCT CONCENTRATIONS

The most common reaction is to add water to an inorganic compound to make a dilute solution of the compound.  
**Hydrogen bond**



**Relative formula mass**  
Is calculated by adding the atomic masses.

**Relative molecular mass**  
Is calculated by adding the atomic masses of all the atoms in the molecule.

**Relative atomic mass**  
Is calculated by adding the atomic masses of all the atoms in the element.

**Relative formula mass**  
Is calculated by adding the atomic masses of all the atoms in the molecule.

**Relative molecular mass**  
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Displacement reactions always involve the most reactive metal reacting with a bonded atom until the least reactive is unbound.

1) Copper sulfate + Magnesium  $\rightarrow$  Magnesium sulfate + Copper

2) Copper sulfate + Zinc  $\rightarrow$  Zinc sulfate + Copper

**Relative formula mass and relative atomic mass**

**Reaction formulae**

$\text{NaCl} + \text{H}_2\text{O} \rightarrow \text{NaOH} + \text{HCl}$

$\text{Na} + 0.020 \text{ g} \quad \text{Cl} + 0.035 \text{ g} \quad \text{H} + 0.001 \text{ g}$

The number of  $\text{Na}^{+}$  atoms and  $\text{Cl}^{-}$  atoms reacting are the same in these ratios.

How to calculate the moles of  $\text{NaOH}$ .

If a reaction is the amounts are present in the same ratio (i.e. 1:1) it reacts.

The excess will be 100% used.

The missing concentration = 0.1 mol/l

**Relative formula mass**  
Is the same as the relative molecular mass.

**Calculate the missing mass of reactant**  
1) 0.005 g  
2) 0.005 g

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Testing for gases

Hydrogen gas - Produces a lighted splint

Oxygen gas - BURNS

Carbon dioxide - Bubbles in lime water

Chlorine gas - Bleaches white paper red

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**Nigel Saunders**

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