

## CH 5 LAB: PERIODIC LAW (CHEM.A.2.3)

**Purpose:** In this experiment, you will use your knowledge of periodic properties and a list of clues to correctly arrange the elements from a scrambled periodic table. You will also predict values for any information missing from the table.

**Theory:** The present organization of the elements is a product of the first periodic table published by Dmitri Mendeleev in 1869. However, the basis of his arrangement was the atomic masses of the elements. This approach proved incorrect, as it would have placed some elements in a family with dissimilar properties. Henry Moseley rearranged the table on the basis of the atomic numbers of the elements. In accordance with Moseley's revision, the Periodic Law states: *the properties of the elements are periodic functions of their atomic numbers*. Each of the 118 known elements has it's own set of characteristic properties. The elements are arranged within the periodic table into groups and periods. This arrangement reflects the periodic or repeating nature of the properties of the elements.

### Materials:

1. Use the following clues to arrange the elements in their proper order in the blank table provided.

The following sets of elements belong together in groups:

- ZRD, PSIF, JXBE, LHT, QKA, WOV, GUN, YMC
- J has an atomic number three times that of T.
- U has a total of six electrons.
- $I_2A$  is the simple formula of an oxide. (Compound with an atom from group 16)
- P is less dense than S.
- S is an alkali metal.
- E is a Noble gas.
- W is a liquid
- Z has the smallest atomic mass in its set.
- B has ten protons.
- O has an atomic number larger than V.
- D has the largest atomic mass of its set.
- C has five electrons in its outer energy level.
- F is a gas.
- X has an atomic number one higher than F.
- L is an alkaline-earth element with atomic mass of 40.
- Y is a metalloid.
- O is a halogen.
- The atomic mass of T is more than that of H.
- Q has an atomic mass 2 times that of A.
- Atoms of I are larger than those of S.
- M has an atomic number one less than that of A.
- The electrons of atom N are distributed over three energy level.
- The atomic radius of K is the largest of the set.

2. Some information is missing from each block. Use your knowledge of periodic properties, and from the information below, to predict the values for the missing items from the location of the element on the periodic table. Place your predictions on your table in periods 5 and 6.

**Data:** Attach Periodic Table on next clean page in notebook.

**\*Conclusion:** In conclusion, the location of an element on the periodic table is based on it's \_\_\_\_\_.

# Periodic Table Clues Lab

**Thomas Paul Hunt**



## **Periodic Table Clues Lab:**

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**Laboratory Operations and Management** Jeanne Isabel, Patricia Ellinger, Dana Griffin, 2024-10-01 Management and lab operations for career entry level students The path to leadership success in the lab begins with a firm foundation in the many facets of management Designed to meet the needs of today s students this new text introduces the principles of laboratory management and demonstrates how to apply them in the real world It builds a foundation in the administration of personnel and finance It also addresses the broad range of knowledge needed to become an effective leader in today s increasingly complex environment quality measures compliance safety regulations and professionalism information critical for all laboratory personnel including the new supervisor **The S Files** Rebecca Linam, 2015-04-25 Warning What you are about to read is a complete and truthful account of the world of espionage Forget everything you thought you knew about spies What you thought you knew is totally wrong For example did you know that the Russian Spy Agency employed seven year old spies the American Spy Agency s top scientist Dr Pepper cloned its most successful agents the Japanese Spy Agency developed a type of cheese that boosted its agents success rates by up to 25% the Spanish Spy Agency invented its famous tickle torture the Flemish Spy Agency uncovered a Martian plot to take over the world the Phlegmish Spy Agency was responsible for many major flu outbreaks during the Cold War You might also think that backwards Russian pig Latin is not a real spy code but it s listed in the Russian Spy Agency handbook This book explains how it really happens This is spying in its true exposed form This book will self destruct when you finish reading it [The Periodic Table](#) Primo Levi, 1984 An extraordinary work in which each of the 21 chapters takes its title and starting point from one of the elements in the periodic table Mingling fact and fiction history and anecdote Levi uses his training as a chemist and his experiences as a prisoner in Auschwitz to illuminate the human condition From the Hardcover edition Copyright Libri GmbH All rights reserved

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**Chemical Building Blocks** Michael J. Padilla, David V. Frank, John G. Little, Steve Miller, 2000

**Virtual Immersive and 3D Learning Spaces: Emerging Technologies and Trends** Hai-Jew, Shalin, 2010-09-30

Virtual Immersive and 3D Learning Spaces Emerging Technologies helps push the conceptual and applied boundaries of virtual immersive learning Virtual immersive spaces bring with them plenty of promise of sensory information rich learning experiences that will enable a much wider range of experiential learning and training delivered to computer desktops augmented reality spaces digital installations and mobile projective devices This work explains how these spaces may be exploited for effective learning in terms of the technologies pedagogical strategies and directions

Spotlight Science Keith Johnson, Sue Adamson, Gareth Williams, 2002 This Spiral Edition Teacher Support Pack offers comprehensive support and guidance providing the best possible learning experience for your students and saving time for everyone in the department

Science Scope, 1998

**The Computing Teacher**, 1989

Chemistry in the Laboratory James M. Postma, Julian L. Roberts, J. Leland Hollenberg, 2009-08-07 For nearly 40 years Chemistry in the Laboratory has been meeting the needs of teachers and students This new edition builds on that legacy while addressing cutting edge trends in the chemistry laboratory including forensic chemistry and environmental and green chemistry As always the new edition of Chemistry in the Laboratory offers precise easy to follow instructions helpful illustrations and an emphasis throughout on laboratory safety Again throughout a Consider This feature encourages users to expand the principles of the experiment into interesting applications open ended experiments or unexplored corners Most experiments in the manual can be completed in one lab session but some can be linked or extended for a multi lab project

**Practical Guide to Computers in Education** Peter Coburn, 1985

Chemistry: Science of Matter, Energy, and Change Gregory R. Choppin, Bernard Jaffe, 1965

**Building**

**Block** Prentice-Hall Staff,1994      *Contributions from the Physics Laboratory* University of Michigan. Department of Physics,1950      **Chemistry and Human Affairs** William Evans Price,George Howard Bruce,1949      **Nelson Textbook of Pediatrics E-Book** Robert Kliegman,Bonita F. Stanton,Joseph W. St. Geme III,Nina F Schor,Richard E. Behrman,2011-06-10

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