

Periodic Trends Worksheet

Using the periodic table and your knowledge of periodic trends, answer the following questions:

1. Circle the best answer that has the:

- a. smallest electron affinity
- b. largest electronegativity
- c. largest ionization energy
- d. lowest ionization energy
- e. smallest atomic radius
- f. largest atomic radius

| | | | | |
|----|----|----|----|----|
| N | P | As | Sb | Bi |
| Ga | Ge | As | Sc | |
| O | S | Sc | Te | Po |
| O | S | Se | Te | Po |
| F | Cl | Br | I | At |
| Li | Na | K | Rb | Cs |

2. Low ionization energy is a characteristic of a _____

3. What does atomic radius measure? _____

4. Describe the change in the atomic size across a period and down a group.

5. Which element has the largest atomic radius? _____

6. Explain the main difference between metals and nonmetals.

7. Determine if the listed pairs of elements have similar or different properties.

- a. Potassium & Rubidium
- c. Sodium & Chlorine
- e. Phosphorus & Oxygen

- b. Calcium & Barium
- d. Helium & Krypton
- f. Lithium & Fluorine

8. Which element has the largest electronegativity? _____

9. Which group (name) would have the lowest ionization energies?

10. Which group (name) has the highest ionization energies?

Periodic Trends Practice Questions Key

Jiyuan Zhang

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