

Periodic Trends Worksheet

Use the periodic table and your knowledge of periodic trends to answer the following questions.

1. Which atom in each pair has the larger atomic radius?
a) Li or **K** b) **Ca** or Ni c) **Ga** or B d) O or **C** e) Cl or **Br**
f) Be or **Ba** g) **Si** or S h) Fe or **Au**

2. What is the periodic trend for atomic size from top to bottom in a group? from left to right in a period?

Top to bottom: increases due to adding energy levels Left to Right: decreases

3. Why do atoms get smaller as you move left to right in a period?
due to increased nuclear charge and no added shielding.

4. Which element in each pair has a larger ionization energy?
a) Na or **O** b) **Be** or Ba c) **Ar** or F d) **Cu** or Ra e) I or **Ne**
f) K or **V** g) **Ca** or Fr h) W or **Se**

5. Explain the relationship between the relative size of an ion to its neutral atom and the charge on the ions.

Positive ions are smaller than the neutral atom it comes from, negative ions are larger than the neutral atom it comes from

6. Which particle has the larger radius in each atom/ion pair?
a) Na, Na^+ b) S, S^{2-} c) I, I^- d) Al, Al^{3+}

7. What is ionization energy? What is first ionization energy?

The amount of energy needed to remove an electron from an atom. First ionization energy is the energy required to remove a single valence electron from an atom.

8. What is the periodic trend for first ionization energy?

It increases going from left to right due to increased nuclear charge, and decreases going down a group due to increased distance from the nucleus and increased shielding.

9. Arrange the following groups of elements in order of increasing ionization energy.

a) Be, Mg, Sr b) Bi, Cs, Ba c) Na, Al, S
high, middle, low high, low, middle low, middle, high

10. Which element in each pair has a higher electronegativity value?

a) Cl, **F** b) C, **N** c) **Mg**, Ne d) **As**, Ca

11. What is the periodic trend for electronegativity? Increases going left to right, decreases going down.

Periodic Trends Questions And Answers

Williams & Wilkins Review



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Cracking the AP Chemistry, 2002-2003 Edition Paul Foglino, Princeton Review (Firm), 2002-01-29 The Princeton Review realizes that acing the AP Chemistry exam is very different from getting straight As in school The Princeton Review doesn't try to teach students everything there is to know about chemistry only the techniques they'll need to score higher on the exam There's a big difference In *Cracking the AP Chemistry* TPR will teach test takers how to think like the test makers and Score higher by reviewing key chemistry concepts Earn more points by becoming familiar with the format of the test Safeguard against traps that can lower scores Perfect skills with review questions in each chapter This book includes 2 full length simulated AP Chemistry exams All of The Princeton Review practice test questions are like the ones test takers will see on the actual exam and every solution is fully explained

Instructor's manual and solutions to problems to accompany General chemistry: principles and structures James E. Brady, 1975-04

Cycles, 1958

Chemistry James E. Brady, John R. Holum, 1995-12-29 Offers accurate lucid and interesting explanations of basic concepts and facts of chemistry while helping readers develop skills in analytical thinking and problems solving

Business Review, 2002

Chemistry & Chemical Reactivity John C. Kotz, Paul Treichel, Patrick A. Harman, 2003 Provides a broad overview of the principles of chemistry the reactivity of chemical elements and their compounds and the applications of chemistry Conveys a sense of chemistry as a field that not only has a lively history but also one that is currently dynamic with important new developments on the horizon

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