

Student Exploration: Electron Configuration

Activity B:
Atomic radii

Get the Gizmo ready:

- On the PERIODIC TABLE tab, select Na (sodium).
- Select the ELECTRON CONFIGURATION tab.



Question: How do the radii of atoms change across a period of the periodic table?

1. **Predict:** Positively charged protons in the nucleus of an atom are attracted to negatively charged electrons.

How do you think the **atomic radii** will change as electrons are added to a shell?

It will decrease

2. **Arrange:** Create a proper electron configuration for sodium. After clicking **Check**, note the **Electron configuration** and the **Atomic radius** now listed at right.

Sodium electron configuration: **1s₂ 2s₂ 2p₆ 3s₁** Atomic radius: **190 Pico meters**

3. **Compare:** Click **Next element**, and then add an electron to the magnesium atom. Click **check**, and record the electron configuration and atomic radius below.

Magnesium electron configuration: **1s₂ 2s₂ 2p₆ 3s₂** Atomic radius: **145 Pico meters**

4. **Gather data:** Create electron configurations for the next six elements. Record the electron configuration and atomic radius of each. (Note: The symbol for picometer is **pm**.)

Element	Number of electrons	Electron configuration	Atomic radius (pm)
Aluminum	13	1s₂ 2s₂ 2p₆ 3s₂ 3p₁	118
Silicon	14	1s₂ 2s₂ 2p₆ 3s₂ 3p₂	111
Phosphorus	15	1s₂ 2s₂ 2p₆ 3s₂ 3p₃	98
Sulfur	16	1s₂ 2s₂ 2p₆ 3s₂ 3p₄	88
Chlorine	17	1s₂ 2s₂ 2p₆ 3s₂ 3p₅	79
Argon	18	1s₂ 2s₂ 2p₆ 3s₂ 3p₆	71

Periodic Table Exploration Answer Key

Ying-Ying Zheng

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Office of Exploration, Exploration Studies Technical Report, FY 1988 Status: Study approach and results ,1988 *Office of Exploration: Exploration Studies Technical Report. Volume 2: Studies Approach and Results ,1988*

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