

Arrange the following elements in order of increasing first ionization energy:

1: Ne, Li, C, N, Be, F = _____

2: Cl, Si, Na, Ar, P, Mg = _____

3: Sr, I, Mo, Cd, Xe, Sb = _____

4: Ba, W, Pb, Cs, Po, Rn = _____

5: Lv, Fl, Rg, Rf, Ra, Fr = _____

6: C, Si, Sn, Ge = _____

7: Li, Rb, K, Na = _____

8: I, Cl, Br, F = _____

9: Nb, Db, V, Ta = _____

10: Ni, Pd, Pt, Ds = _____

— **Electronegativity** is a measure of the tendency of an atom to attract a bonding pair of electrons to itself.

— Moving from left to right across a period electronegativity increases until we reach G17, then it begins to drop rapidly at G18.

— Moving across a period adds electrons to the outer shell of an atom moving it closer to having a full outer orbital, thus increasing its ability to attract electrons; noble gases already have a filled outer orbital and have low electronegativity.

— Moving from the top to the bottom of a group, the electronegativity decreases.

— Moving down a group adds energy levels filled with electrons that shield protons from attracting outer shell electrons.

Arrange the following elements in order of increasing electronegativity energy:

11: Li, B, N, Be, F = _____

12: Cl, Al, Na, P, S = _____

13: Sr, I, Te, Pd, Sn = _____

14: Ra, Ag, Ta, Ge, S = _____

15: Lv, Fl, Rg, Rf, Fr = _____

16: C, Si, Sn, Ge = _____

17: Li, Rb, K, Na = _____

18: I, Cl, Br, F = _____

19: Mo, W, Sg, Cr = _____

20: Ni, Pd, Pt, Ds = _____

The **melting point** of an element is defined as the temperature at which a solid changes phase to a liquid.

— Moving left to right across a period, G1 elements have low melting points and the melting point increases to its maximum value between G6 and G14, then drops back to a low value at G18.

— Elements in G1 and G2 are gases or low molecular weight elements with s-orbital valance electrons.

The transition metals elements in the middle of the table have higher melting points. All periods end with the noble gases and like all gases G18 elements have very low melting points.

— Moving down a group, G1 and G2 elements the melting point of the elements remains fairly constant.

Periodicity Lab Answer

Judd E. Hollander



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