

# Periodic Table (Advanced)

The periodic table is a particular arrangement of the chemical elements, organized on the basis of their atomic numbers (number of protons in the nucleus), electron configurations, and resulting chemical properties. Elements are presented in order of increasing atomic number, which is typically listed with the element symbol in each box. The standard form of the table consists of a grid of elements laid out in 18 columns and 7 rows, with a similar set of elements below that. The table contains the above numbered sets from left to right: the 18-block to the left, the 16-block to the right, the 12-block in the middle, and the 16-block below that.

The rows of the table are called periods, the columns are called groups, with some of these having further sub-blocks as indicated in the grid. Thus, by definition, a periodic table incorporating existing trends, any such table can be used to determine relationships between the properties of the elements and predict the properties of new, yet to be discovered or synthesized, elements. As a result, a periodic table—whether in the standard form or some other format—provides a useful framework for analyzing chemical behavior, and such tables are widely used in chemistry and other sciences.

Although present under Dmitri Mendeleev's name, the periodic table is generally credited with the publication of the first, and still widely recognized periodic table, the developed by Mendeleev to illustrate periodic trends in the properties of the then-known elements. Mendeleev also predicted some properties of then-unknown elements that would be expected to fall gaps in this table. Most of his predictions were proved correct when the elements he predicted were subsequently discovered. Mendeleev's periodic table has since been expanded and refined with the discovery or synthesis of further new elements and the development of new theoretical models to explain chemical behavior.

All elements from atomic number 1 (hydrogen) to 118 (oganesson) have been discovered or reportedly synthesized, with others, 119, 120, 121, 122, and 123, yet to be confirmed. The first 98 elements exist naturally although some are found only in trace amounts and were synthesized in laboratories before being found in nature. Elements 99 through 104 are synthetic elements from 197 to 200 that may have been synthesized, or claimed to be so, in laboratories. Production of elements having higher atomic numbers is being pursued, with the question of how the periodic table may need to be modified to accommodate very much additional (being a measure of ongoing studies). These new synthetic elements are often naturally occurring elements that also have yet to be discovered.

All groups of the periodic table include some chemical elements, but electrons, neutrons, protons, or molecular particles. Each element of a group has a unique atomic number representing the position of persons in the nucleus. Most elements have a single isotope, or isotope, among them, with those isotopes being referred to as isotopes. For example, carbon has three naturally occurring isotopes, all of its atoms have six protons and most have six neutrons as well, but about one percent have seven neutrons, and a very small fraction have eight neutrons. Isotopes are never separated in the periodic table, they are always grouped together under a single element. Elements with two stable isotopes have their atomic masses of their most stable isotopes, with much smaller numbers, listed in parentheses.

In the standard periodic table, the elements are listed in order of increasing atomic number (the number of persons in the nucleus of an atom). A note: some (particularly in standard) tables also include atomic mass but this is not always. Groups 1 and 2 are determined by the electron configurations of the atoms elements with the same number of electrons in a particular shell (that is to the same column). Groups 13-18, nitrogen and chlorine, are the same column because they have the same number of valence electrons (the outermost or outer shell). Elements with similar chemical properties generally fall into the same groups in the periodic table, although in the 1-block, a few notable exceptions (the alkali metals, the elements in the 18-block) have similar properties, as well. Thus, it is relatively easy to predict the chemical properties of an element if one knows the properties of the elements in group 1.

As of 2023, the periodic table has 114 confirmed elements, ranging from hydrogen (symbol H) (hydrogen) to 118 (oganesson), 119 (darmstadtium) and 123 (moscovium). Elements 113, 115, 117, and 118 are, respectively the 18th, 19th, 20th, and 21st elements in the International Union of Pure and Applied Chemistry (IUPAC). As such, these elements are currently known only by their systematic element names, listed in their atomic numbers.

A total of 187 elements occur naturally—the remaining 16 elements, from plutonium to neptunium, and the rest are man-made elements, most only when synthesized in laboratories. Of the 187 elements that occur naturally, 94 are primordial. The other 93, naturally occurring elements occur only as decay chains of primordial elements. One element between these elements (technetium) has never been observed in existence per se, only in the form of.



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*Advances in Materials Manufacturing Science and Technology II* Cheng Yu Jiang, Geng Liu, Ding Hua Zhang, Xi Peng Xu, 2006-12-15 Selected Papers from the 12th International Manufacturing Conference in China September 21-23 2006 Xi'an China

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