

Periodic Trends

Period

1. Discuss the importance of Mendeleev's periodic law.
When elements are arranged by increasing number of protons (atomic number), there are repeating patterns of chemical and physical properties.
2. Identify each element as a metal, metalloid, or nonmetal.
 - a) Fluorine **Nonmetal**
 - b) Germanium **Metalloid**
 - c) zinc **Metal**
 - d) phosphorous **Nonmetal**
 - e) lithium **Metal**
3. Give two examples of elements for each category.
 - a) noble gases **He, Ne, Ar, Kr, Xe, Rn, 118**
 - b) halogens **F, Cl, Br, I, At, 117**
 - c) alkali metals **Li, Na, K, Rb, Cs, Fr**
 - d) alkaline earth metals **Be, Mg, Ca, Sr, Ba, Ra**
3. What trend in atomic radius do you see as you go down a group/family on the periodic table? What causes this trend? **Atomic radii INCREASE going down a group due to added energy levels and increases shielding effects.**
4. What trend in atomic radius do you see as you go across a period/row on the periodic table? What causes this trend? **Atomic radii DECREASE going across a period due to added energy levels and increases shielding effects.**
5. Circle the atom in each pair that has the largest atomic radius.
 - a) **Al** B d) Na **Al**
 - b) **S** O e) O **F**
 - c) **Br** Cl f) Mg **Ca**
6. Define ionization energy.
The amount of energy required to remove a valence electron.
7. What trend in ionization energy do you see as you go down a group/family on the periodic table? What causes this trend? **Ionization energy DECREASES as you go down a group, as the atomic radius increases. This is due to greater shielding effects.**

Periodic Trends Quiz Answer Key

David Kirk

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Table of Contents Periodic Trends Quiz Answer Key

1. Understanding the eBook Periodic Trends Quiz Answer Key
 - The Rise of Digital Reading Periodic Trends Quiz Answer Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Periodic Trends Quiz Answer Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Periodic Trends Quiz Answer Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from Periodic Trends Quiz Answer Key
 - Personalized Recommendations
 - Periodic Trends Quiz Answer Key User Reviews and Ratings
 - Periodic Trends Quiz Answer Key and Bestseller Lists
5. Accessing Periodic Trends Quiz Answer Key Free and Paid eBooks

- Periodic Trends Quiz Answer Key Public Domain eBooks
- Periodic Trends Quiz Answer Key eBook Subscription Services
- Periodic Trends Quiz Answer Key Budget-Friendly Options
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 - ePub, PDF, MOBI, and More
 - Periodic Trends Quiz Answer Key Compatibility with Devices
 - Periodic Trends Quiz Answer Key Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Periodic Trends Quiz Answer Key
 - Highlighting and Note-Taking Periodic Trends Quiz Answer Key
 - Interactive Elements Periodic Trends Quiz Answer Key
- 8. Staying Engaged with Periodic Trends Quiz Answer Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Periodic Trends Quiz Answer Key
- 9. Balancing eBooks and Physical Books Periodic Trends Quiz Answer Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Periodic Trends Quiz Answer Key
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Periodic Trends Quiz Answer Key
 - Setting Reading Goals Periodic Trends Quiz Answer Key
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Periodic Trends Quiz Answer Key
 - Fact-Checking eBook Content of Periodic Trends Quiz Answer Key
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development

- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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